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 $q = mc\Delta T = 1.00 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 10.5^\circ\text{C}$
 $q = 43.9 \text{ J}$
 Statement:
 The amount of energy absorbed by 1.00 g of water, H₂O(l), when its temperature is raised 10.5 °C is 43.9 J. 36.
 Given: V
 HCl(aq) = 100.0 mL; V
 NaOH(aq) = 100.0 mL; T
 initial = 23.5 °C; T